

Name: \_\_\_\_\_

Period: \_\_\_\_\_ Table: \_\_\_\_\_

Fill-in each block with the correct data for the compounds noted. **Show your Work!**

Compound		A			B		
		# of Valence Electrons for each element	Draw Individual Lewis Dot Diagrams	Draw Combined Lewis Dot Diagrams	Individual Electronegativity	Electronegativity Difference (end)	Type of Bond
1	Example	Be = 2 Br = 7	<div>Be•</div> <div>•••••</div> <div>•••••</div> <div>•••••</div>	<div>Be•••••</div> <div>•••••</div> <div>•••••</div> <div>•••••</div>	Be = 1.5 Br = 2.8	end = 2.8 - 1.5 end = 1.3	Polar Covalent
2							
	H <sub>2</sub> O						
3							
	NaCl						
4							
	CH <sub>4</sub>						

Compound		# of Valence Electrons for each element	Draw Individual Lewis Dot Structures	Draw Combined Lewis Dot Structures	Individual Electronegativity	Electronegativity Difference	Type of Bond
5							
	Br <sub>2</sub>						
6							
	NH <sub>3</sub>						
7							
	MgCl <sub>2</sub>						
8							
	KF						
9							
	Li <sub>2</sub> O						
10							
	CF <sub>4</sub>						